

Please replace the paragraph bridging pages 10 and 11 (line 23, page 10 through line 10, page 11), with the following rewritten paragraph:

According to this [[Carey-Lee]] Carey-Lea process, an aqueous silver nitrate solution is mixed in a mixed solution of an aqueous iron(II) sulfate solution and an aqueous sodium citrate solution to allow them to react, the resultant agglomerate of fine silver particles are filtered and washed, and thereafter pure water is added to the resultant cake of the agglomerate of fine silver particles, whereby a fine silver particle colloidal dispersion can be obtained. The fine silver particle colloidal dispersion obtained by this Carey-Lee process has a fine silver particle concentration of from 0.1 to 10% by weight. If it has a concentration of more than 10% by weight, it tends to agglomerate.

OK 6/29/10
Please replace the paragraph bridging pages ²¹22 and ²²23 (line 16, page ²¹22 through line 19, page ²²23), with the following rewritten paragraph:

As the organic solvent, it is preferable to use an organic solvent containing at least dimethyl sulfoxide. This is because the dimethyl sulfoxide acts on the fine silver particles to have the effect of improving the stability of the fine silver particle colloidal dispersion. In the fine silver particle colloidal dispersion of a water-organic solvent system according to the present invention, the fact that the dimethyl sulfoxide has the effect of improving the stability of the fine silver particle colloidal dispersion has been found by the fact that a conductive silver film formed using the fine silver